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Chem chapter 11 vocab Stoichiometry: The study of quantitative relationships between the amounts of reactants used and products formed by a chemical reaction; is based on the law of conservation of mass. mole ratio: In a balanced equation, the ratio between the numbers of moles of any two substances.

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limiting reactant: A reactant that is totally consumed during a chemical reaction, limits the ...

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stoichiometry. mole ratio. limiting reactant. excess reactant. the study of quantitative relationships between the amounts of... in a balanced equation, the ratio between the number of moles... a reactant that is totally consumed during a chemical reaction...

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Key

368 Chapter 11 • Stoichiometry Section
111.11.1 Objectives Describe the types of
relationships indicated by a balanced
chemical equation. State the mole ratios from
a balanced chemical equation. Review
Vocabulary reactant: the starting substance
in a chemical reaction New Vocabulary
stoichiometry mole ratio Defining
Stoichiometry

Chapter 11: Stoichiometry

11.1 Defining Stoichiometry 11.2
Stoichiometric Calculations 11.3 Limiting

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Chemistry Matter and Change: Chapter 11 Stoichiometry ...

In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and

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the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

Chapter 11.4: Stoichiometry - Chemistry

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• Chapter 11 209 StoichiometryStoichiometry

CHAPTER 11 SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372 Practice

Problems pages 371–372 1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show

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that the law of conservation of mass is

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Stoichiometry Chapter 11 Study Guide Answer Key Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

Chapter 11 Stoichiometry Study Answer Key

Chapter 11 187 Exercise 11.3 - Equation
Stoichiometry: Iron is combined with carbon in a series of reactions to form pig iron, which is about 4.3% carbon. $2C + O_2 \rightarrow 2CO$ $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ $2CO + C \rightarrow C$ (in iron) CO_2 Pig iron is

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easier to shape than pure iron, and the presence of carbon lowers its melting point

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Molecules • Number of Moles • Mass o Law of
Conservation of Mass • Volume

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quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

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Section 11.1 What is stoichiometry? 1. true
2. true 3. false 4. true 5. true 6. 2, 2,
64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4,
72.08 10. methanol and oxygen gas 11. carbon
dioxide and water 12. 160.10 g 13. 160.10 g
14. They are equal. 15. A mole ratio is a

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ratio between the numbers of moles

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answer study of quantitative relationships

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